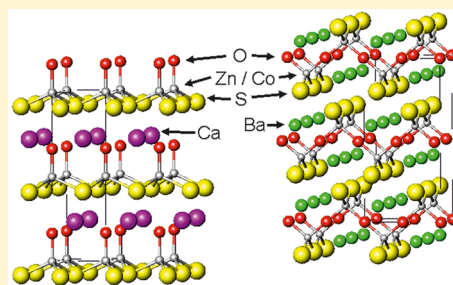


## Crystal and Magnetic Structures of the Oxide Sulfides CaCoSO and BaCoSO

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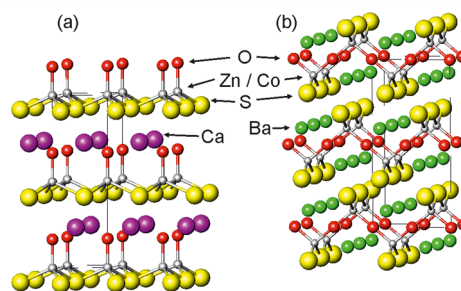
## S Supporting Information

**ABSTRACT:** CaCoSO, synthesized from CaO, Co, and S at 900 °C, is isostructural with CaZnSO and CaFeSO. The structure is non-centrosymmetric by virtue of the arrangement of the vertex-sharing CoS<sub>3</sub>O tetrahedra which are linked by their sulfide vertices to form layers. The crystal structure adopts space group *P6<sub>3</sub>mc* (No. 186), and the lattice parameters are *a* = 3.7524(9) Å and *c* = 11.138(3) Å at room temperature with two formula units in the unit cell. The compound is highly insulating, and powder neutron diffraction measurements reveal long-range antiferromagnetic order with a propagation vector *k* = (1/3, 1/3, 1/2). The magnetic scattering from a powder sample can be modeled starting from a 120° arrangement of Co<sup>2+</sup> spin vectors in the triangular planes and then applying a canting out of the planes which can be modeled in the magnetic space group *C<sub>2</sub>c* (space group 9.40 in the Belov, Neronova, and Smirnova (BNS) scheme) with Co<sup>2+</sup> moments of 2.72(5) μ<sub>B</sub>. The antiferromagnetic structure of the recently reported compound BaCoSO, which has a very different crystal structure from CaCoSO, is also described, and this magnetic structure and the magnitude of the ordered moment (2.75(2) μ<sub>B</sub>) are found by experiment to be similar to those predicted computationally.



## ■ INTRODUCTION

Oxide chalcogenides are a class of compounds in which ordering of the two different anions, O<sup>2-</sup> and the larger and more polarizable S<sup>2-</sup>, Se<sup>2-</sup>, or Te<sup>2-</sup>, results in layered structures with segregation of metal cations according to their chemical preferences. These compounds have been of considerable interest as possible transparent p-type transparent conductors,<sup>1</sup> and the investigation of related mixed-anion compounds led to the unexpected discovery of superconductivity in appropriately substituted lanthanide iron oxide arsenides.<sup>2</sup> In turn, the investigation of the iron-based arsenide and chalcogenide superconductors has led to renewed exploration of oxide chalcogenides. The range of structure types and some of the properties of these compounds have recently been reviewed.<sup>3</sup> We and others reported previously<sup>4–6</sup> the discovery of new structure types for oxide sulfides in the white wide-band gap insulators containing alkaline earth and zinc cations. The non-centrosymmetric polar oxide sulfide CaZnSO with the structure shown in Figure 1a adopts a layered structure in which [ZnS<sub>3</sub>O] tetrahedra share all three basal sulfide vertices with other tetrahedra to form  $\infty$ [ZnS<sub>3</sub>/3O] layers and the tetrahedra are all oriented along the *c* axis resulting in the non-centrosymmetric space group *P6<sub>3</sub>mc*.<sup>4,5</sup> When the alkaline earth cation is the much larger Ba<sup>2+</sup>, a different structure is adopted for BaZnSO (Figure 1b) in which [ZnS<sub>2</sub>O<sub>2</sub>] tetrahedra are linked by sharing each vertex with another tetrahedron.<sup>6</sup> A previous report<sup>7</sup> on the existence of CaFeSO has recently been augmented by further reports characterizing this magnetic Mott–Hubbard insulating compound,<sup>8,9</sup> which is isostructural



**Figure 1.** Crystal structures of (a) CaMSO (*M* = Fe,<sup>8,9</sup> Co, Zn<sup>4,5</sup>) and (b) BaMSO (*M* = Co,<sup>10</sup> Zn<sup>6</sup>).

with CaZnSO. BaCoSO,<sup>10</sup> isostructural with BaZnSO, has also been described. Here we characterize a third member with the CaZnSO structure, CaCoSO (Figure 3), and we make comparisons of the crystal structures and the magnetic ordering in this series of compounds. We also report the magnetic structure of BaCoSO and show that it is as predicted by recent computational investigations.<sup>10</sup>

## ■ EXPERIMENTAL SECTION

**Synthesis.** Because of the air sensitivity of some of the reactants, solids were handled in an argon-filled glovebox. Both compounds were synthesized from stoichiometric amounts of alkaline earth oxides and elemental Co and S with the ground mixtures contained in alumina

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crucibles within dried evacuated silica tubes. In both cases the initial heating stage was performed at  $2\text{ }^{\circ}\text{C min}^{-1}$  to enable elemental sulfur to react without generating a large overpressure.  $\text{CaCoSO}$  was synthesized on the 4 g scale using  $\text{CaO}$  (obtained from the thermal decomposition of  $\text{CaCO}_3$  (ALFA 99.99%)),  $\text{Co}$  (ALFA 99.99%), and  $\text{S}$  (ALFA 99.9995%). The reaction was heated at  $900\text{ }^{\circ}\text{C}$  for two periods of 2 days with an intermediate regrinding, and this resulted in material that was free of impurities as judged by laboratory X-ray powder diffraction (XRPD). Lower reaction temperatures resulted in impure products.  $\text{BaCoSO}$  was synthesized on a similar scale according to the method described in ref 10.  $\text{BaO}$  (Aldrich 99.99%),  $\text{Co}$  (Alfa Aesar 99.998%), and  $\text{S}$  (Alfa Aesar 99.999%) were ground together, pressed into a pellet (300 MPa), and heated at  $900\text{ }^{\circ}\text{C}$  for 12 h (after heating at  $2\text{ }^{\circ}\text{C min}^{-1}$ ). This product was reground, repelletized, and heated at  $940\text{ }^{\circ}\text{C}$  for 100 h.

**Diffraction Measurements.** Laboratory XRPD patterns to monitor the progress of reactions and for initial structural characterization were measured using a PANalytical X'Pert diffractometer with  $\text{Cu K}\alpha_1/\alpha_2$  radiation or with a PANalytical Empyrean instrument with  $\text{Cu K}\alpha_1$  radiation. Neutron powder diffraction (NPD) data for  $\text{CaCoSO}$  were collected at ambient temperature and at 8 K on a 4 g sample using the now-retired diffractometer D1A at the ILL Grenoble. Data for  $\text{BaCoSO}$  were collected at ambient temperature and 3.5 K using a 4.5 g sample using the diffractometer D2B at ILL. The samples were contained in thin-walled vanadium cylinders. For the low temperature measurements the samples were cooled using different closed-cycle refrigerators. Refinements against the diffraction data were carried out using Topas Academic V5,<sup>11</sup> and magnetic scattering was modeled by introducing additional magnetic-only phases to the refinements. Magnetic models were deduced with the aid of the ISODISTORT software suite.<sup>12</sup>

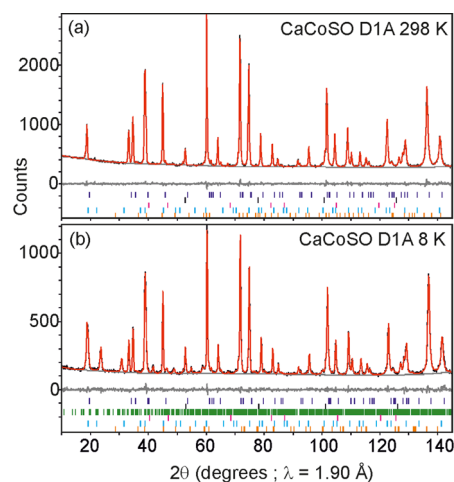
**Physical Property Measurements.** Magnetic susceptibility measurements for  $\text{CaCoSO}$  were made using a Quantum Design MPMS-5 SQUID magnetometer. 4-Probe resistivity measurements on a sintered pellet of  $\text{CaCoSO}$  were attempted using a home-built apparatus.

## RESULTS AND DISCUSSION

### Structure, Properties, and Magnetic Order of $\text{CaCoSO}$ .

Analysis of the XRPD pattern of  $\text{CaCoSO}$  revealed that the phase was not contaminated by significant amounts of impurity phases and that the phase was isostructural with  $\text{CaZnSO}$ . The structure at ambient temperature was refined from the NPD pattern (Figure 2a), and small impurities, namely,  $\text{CaO}$  (0.4% by mass),  $\text{Co}_9\text{S}_8$  (3.2%), and  $\text{CaCO}_3$  (calcite; 0.2%), were included in the refinement along with reflections from the vanadium sample holder. The refinement results at room temperature are presented in Table 1 and Table 2. Table 3 summarizes the key structural parameters for the three isostructural compounds  $\text{CaZnSO}$ ,  $\text{CaFeSO}$ , and  $\text{CaCoSO}$ .

Comparison of bond lengths and angles in the three compounds reveals a trend in the  $\text{M}-\text{O}$  and  $\text{M}-\text{S}$  distances from  $\text{M} = \text{Fe}$  to  $\text{Co}$  which reflects the normal trends across the series. In  $\text{CaZnSO}$  the ratio of the  $\text{Zn}-\text{S}$  and  $\text{Zn}-\text{O}$  distances is similar to that calculated from the tables of published ionic radii,<sup>13</sup> while for both  $\text{CaFeSO}$  and  $\text{CaCoSO}$  the  $\text{M}-\text{O}$  distance is relatively short, consistent with relatively strong covalent bonding between the transition metal cations and oxide. A comparison of the  $\text{Co}-\text{O}$  and  $\text{Co}-\text{S}$  distances in  $\text{CaCoSO}$  with the values obtained for spinels, where  $\text{Co}^{2+}$  ions occupy regular tetrahedral sites, reveals that in  $\text{CaCoSO}$  the  $\text{Co}-\text{O}$  distance is 4.3% shorter than that in  $\text{Co}_3\text{O}_4$ <sup>14</sup> and the  $\text{Co}-\text{S}$  distance is 1.7% longer than that in  $\text{Cr}_2\text{CoS}_4$ .<sup>15</sup> In  $\text{BaCoO}_2$ , one of few representatives of oxides containing tetrahedral  $\text{Co}^{2+}$  ions and an electropositive metal cation,<sup>16</sup> the mean  $\text{Co}-\text{O}$  distance of  $1.904\text{ }\text{\AA}$  is 2.4% longer than in



**Figure 2.** Rietveld refinement against NPD data for  $\text{CaCoSO}$  at (a) 298 K and (b) 8 K. Data (black line), calculated (red), difference (gray) are shown, and the background function (gray line) is also shown. Tick marks show reflection positions. From top to bottom these are for the  $\text{CaCoSO}$  nuclear phase (purple), vanadium sample holder (black),  $\text{CaCoSO}$  magnetic phase (green, 8 K data only),  $\text{CaO}$  (0.4% by mass) (magenta),  $\text{Co}_9\text{S}_8$  (3.2%) (cyan), and  $\text{CaCO}_3$  (calcite; 0.2%) (orange).

**Table 1. Refinement Results for  $\text{CaCoSO}$  at Room Temperature**

formula	$\text{CaCoSO}$
radiation	neutrons $\lambda = 1.901\text{ }\text{\AA}$
instrument	D1A (ILL)
physical form	black powder
$T$ (K)	298
cryst syst	hexagonal
space group	$P6_3mc$ (No. 186)
formula wt	147.1
$a$ ( $\text{\AA}$ )	3.7415(8)
$c$ ( $\text{\AA}$ )	11.106(2)
$V$ ( $\text{\AA}^3$ )	134.64(6)
$Z$	2
$\rho_{\text{calc}}$ ( $\text{Mg m}^{-3}$ )	3.628(2)
$\chi^2$	1.23
$wR_p$	0.0255
$R_{\text{Bragg}}$	0.0083

**Table 2. Refined Atomic Parameters at Room Temperature for  $\text{CaCoSO}$**

atom	site	$x$	$y$	$z$	$100 \times (U_{\text{iso}} (\text{\AA}^2))$	Occ <sup>a</sup>
Ca	2b	1/3	2/3	0.2682(3)	0.59(5)	1
Co	2a	0	0	0 <sup>b</sup>	0.66(8)	1
S	2b	2/3	1/3	0.0789(4)	0.50(8)	1
O	2a	0	0	0.3326(3)	0.66(3)	1

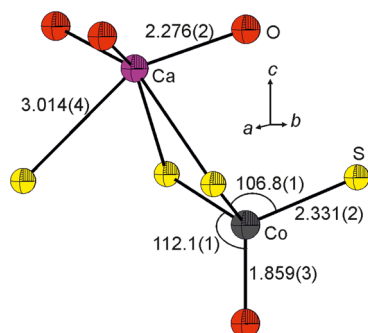
<sup>a</sup>Not refined. <sup>b</sup> $z$  coordinate for Co fixed at 0 in this non-centrosymmetric space group  $P6_3mc$ .

$\text{CaCoSO}$ , reflecting the relative shortness of the  $\text{Co}-\text{O}$  distance in  $\text{CaCoSO}$ . The data also reveal that, of the three compounds, the angles in the  $\text{CoS}_3\text{O}$  tetrahedron are the closest to the angles in a regular tetrahedron. The ligand field in  $\text{CaCoSO}$  is discussed below.

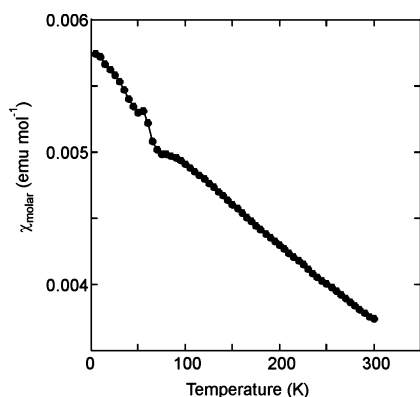
The magnetic susceptibility of  $\text{CaCoSO}$  is shown in Figure 4. A measurement of magnetization against applied field at room

**Table 3.** Selected Bond Lengths (Å) and Angles (deg) for CaCoSO Compared with Literature Values for CaFeSO<sup>8</sup> and CaZnSO<sup>5</sup>

bond or angle	CaFeSO <sup>8</sup>	CaCoSO	CaZnSO <sup>5</sup>
Ca–O [3] (Å)	2.298(4)	2.276(2)	2.2857(3)
Ca–S [3] (Å)	2.982(9)	3.014(4)	3.0346(8)
M–O [1] (Å)	1.88(1)	1.859(3)	1.8996(5)
M–S [3] (Å)	2.391(6)	2.331(2)	2.3716(5)
M–S/M–O	1.27(1)	1.254(3)	1.248(1)
O–M–S [3] (deg)	114.76(4)	112.1(1)	113.84(3)
S–M–S [3] (deg)	103.70(3)	106.8(1)	104.77(3)



**Figure 3.** Structure of CaCoSO showing isotropic displacement ellipsoids (at the 99% level) and selected bond lengths (in Å) and angles (in degrees).

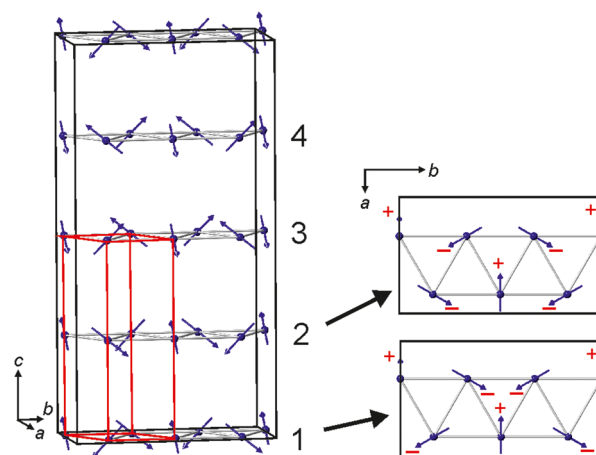


**Figure 4.** Magnetic susceptibility of CaCoSO as a function of temperature.

temperature was nonlinear, suggesting that a minuscule (below the detection limits of the diffraction experiments) ferromagnetic impurity was present. The magnetic susceptibility was therefore obtained by making measurements as a function of temperature at 3 and 4 T, fields above the saturation field of the impurity, and obtaining the temperature dependence of the susceptibility from the differences between the values measured at the two fields. The sample was cooled in zero applied field. The magnetic susceptibility reveals a kink at around 80 K, which we presume indicates the onset of magnetic order, and there is a further feature at around 60 K. The sample did not obey the Curie–Weiss law at high temperatures: a plot of the inverse susceptibility against temperature in the range 200–300 K superficially appeared linear, but the intercept with the temperature axis (i.e., the value of the Weiss temperature) was about  $-500$  K, so any attempt to extract a reliable effective magnetic moment for  $\text{Co}^{2+}$  was not justified. The sample was found to be highly insulating, and the resistance of the pellet

was too high to enable a value for the resistivity to be extracted. This is consistent with localization of  $\text{Co}^{2+}$  electrons.

At 8 K new reflections are evident in the NPD pattern of CaCoSO (Figure 2b) which are consistent with magnetic ordering. The positions of the magnetic reflections are clearly different from those reported for CaFeSO.<sup>9</sup> Analysis of these new reflections using the ISODISTORT software<sup>12</sup> revealed that their positions could be accounted for by a magnetic structure with a propagation vector,  $k = (1/3, 1/3, 1/2)$ . Analysis of possible distortion modes that would account for the intensities of these reflections showed that most of the magnetic scattering could be accounted for by arranging the  $\text{Co}^{2+}$  moments, which describe a triangular lattice in the  $ab$  plane, on a  $120^\circ$  structure with the moments oriented within the triangular plane. This is one way to satisfy the inherent frustration between antiferromagnetic alignment of nearest neighbor spins and their arrangement on a triangular lattice. This model required a single parameter corresponding to the magnitude of the  $\text{Co}^{2+}$  moment. However, some discrepancies remained in the intensities of the most intense magnetic reflections (see Figure S1). Introduction of an out-of-plane canting distortion of the moments enabled these discrepancies to be accounted for by the introduction of one further refined parameter to model the canting and with constraints applied to ensure that the  $\text{Co}^{2+}$  ions all carried equal moments. The intensity discrepancies were only accounted for by a canting in which the moments within a single triangular Co layer were not all canted in the same sense. This resulted in the loss of 3-fold rotational symmetry, and so the magnetic model shown in Figure 5 is monoclinic and is described in the magnetic space



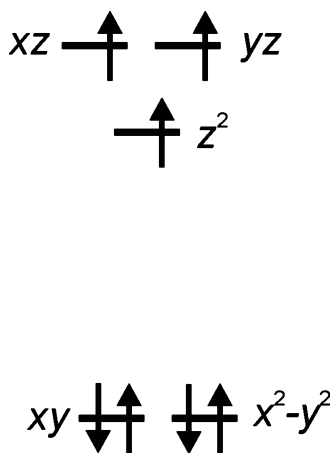
**Figure 5.** Magnetic model for CaCoSO derived from NPD data. The dimensions of the structural unit cell are depicted in red. Co atoms are shown. Other atoms are omitted for clarity. Tables S4 and S5 describe the model in detail. At right are shown the arrangements of Co moments in the layers labeled 1 and 2 viewed perpendicular to the layers. The canting (moments labeled “+” are canted to point above the plane; moments labeled “−” are canted to point below the plane) destroys the 3-fold rotational symmetry. Layers 3 and 4 are aligned antiferromagnetically with respect to layers 1 and 2 respectively.

group  $C_c$  (space group 9.40 in the Belov, Neronova, and Smirnova (BNS) scheme which is a setting of space group  $C_{2v}m'$  (8.6.43) in the Opechowski and Guccione (OG) scheme).<sup>17</sup> The ordered moment on  $\text{Co}^{2+}$  was refined to be  $2.75(5) \mu_B$ . The fit to the NPD data at 8 K is shown in Figure 2b, and full lists of refinement results are included in Tables



S1–S5). It should be noted that a powder neutron diffraction experiment places intrinsic limitations on the ability to unambiguously determine the relative moment directions,<sup>18</sup> and several models were found which could account equally well for the observed magnetic intensities. These models were all rooted in the 120° structure with an out-of-plane canting applied as described above, and we have chosen a magnetic model in which the moments carried by all the Co<sup>2+</sup> ions have equal magnitudes because the nuclear model contains a single Co<sup>2+</sup> ion. A single crystal neutron diffraction experiment could, in principle, be used to distinguish these models. In the refinements we retained the ambient temperature nuclear model at low temperature as there was no evidence, at the resolution of our instrument, for a structural distortion accompanying the magnetic ordering.

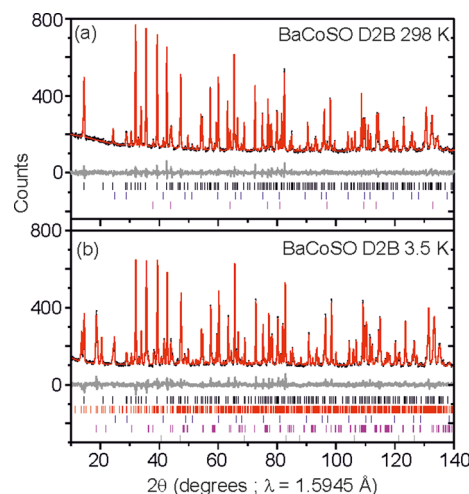
The Co<sup>2+</sup> coordination environment has C<sub>3v</sub> symmetry, and, as described above, the Co–O distance is relatively short and the Co–S bond is relatively long. There are several compounds in the bioinorganic chemistry literature that have a similar distortion of the tetrahedral ligand field around the transition metal ion. An example that mimics well the CaCoSO case is the compound [PhB(CH<sub>2</sub>P(*i*Pr)<sub>2</sub>)<sub>3</sub>]FeN<sup>19</sup> in which there are a strongly  $\sigma$  and  $\pi$  donating apical ligand (nitride) and  $\sigma$ -donor ligands (phosphine) describing the basal triangle. DFT calculations on [PhB(CH<sub>2</sub>P(*i*Pr)<sub>2</sub>)<sub>3</sub>]FeN quantify that in C<sub>3v</sub> symmetry the degenerate d<sub>xz</sub>/d<sub>yz</sub> pair will lie highest in energy as they are  $\pi$ -antibonding with respect to the apical ligand and  $\sigma$ -antibonding with respect to the basal ligands. The degenerate d<sub>xy</sub>/d<sub>x<sup>2</sup>–y<sup>2</sup></sub> pair will lie low in energy. Calculations using a range of similar molecules with this distorted tetrahedral geometry<sup>20–23</sup> show that the energy of d<sub>z<sup>2</sup></sub> is sharply dependent on the geometry and on the bonding characteristics of the ligands. Since in CaCoSO the distortion from tetrahedral is relatively small and is fairly similar to that in [PhB(CH<sub>2</sub>P(*i*Pr)<sub>2</sub>)<sub>3</sub>]FeN,<sup>19</sup> we propose that the occupancy of the d levels will be as calculated for that compound: d<sub>z<sup>2</sup></sub> will be comparable in energy to the higher-lying degenerate d<sub>xz</sub>/d<sub>yz</sub> pair and the configuration will be {(d<sub>xy</sub>)<sup>2</sup>(d<sub>x<sup>2</sup>–y<sup>2</sup>)<sup>2</sup>}{(d<sub>xz</sub>)<sup>1</sup>(d<sub>yz</sub>)<sup>1</sup>}(d<sub>z<sup>2</sup></sub>)<sup>1</sup> (Figure 6). There will be three unpaired electrons and no first order orbital contribution to the moment. The refined value of the moment of 2.72(5)  $\mu_B$  is consistent with this high-spin scenario given that there will be considerable covalency which will lower the value of the ordered moment from the spin-only value of 3  $\mu_B$ .</sub>



**Figure 6.** Proposed ligand field for CaCoSO based on analogous molecular systems.

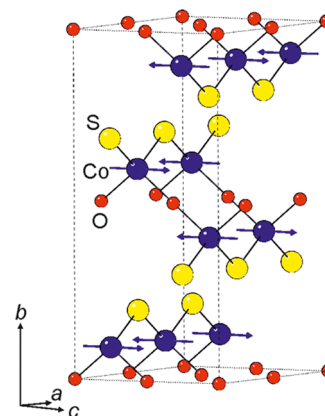
Further measurements on single crystal samples would be required to further examine the exchange interactions between Co<sup>2+</sup> moments to rationalize the magnetic structure in this system with a non-centrosymmetric nuclear structure.

**Magnetic Structure of BaCoSO.** The sample of BaCoSO contained some impurity phases (BaS (4% by mass) and CoO (2.4% by mass)) which were readily accounted for in the refinements. Refinement against D2B data at ambient temperature (Figure 7a, Tables S6 and S7) confirmed the structural



**Figure 7.** Refinements against NPD data for BaCoSO. (a) Room temperature (nuclear model only). (b) 3.5 K (nuclear and magnetic models). The refinements confirm the nuclear structure found by Valldor et al.<sup>10</sup> and their magnetic structure proposed by computational methods. Full refinement details are given as Supporting Information. The figures show the data (black line), fit (red line), and difference (gray line) together with tick marks for the phases present. From top: BaCoSO nuclear (black), BaCoSO magnetic (red, 3.5 K data only), BaS (4% by mass, blue), CoO (2.4%, pink), Al sample environment (gray, 3.5 K data only).

model reported by Valldor et al. in space group *Cmcm*.<sup>10</sup> At 3.5 K additional reflections were evident (Figure 7b) and the proposed antiferromagnetic model with the lowest computed energy reported by Valldor et al.<sup>10</sup> and shown in Figure 8,



**Figure 8.** Magnetic structure of BaCoSO showing the antiferromagnetic arrangement of nearest neighbor moments via Co–O–Co and Co–S–Co linkages. Ba atoms are omitted for clarity. The unit cell vectors are depicted for the magnetic cell (see Table 4), which has double the volume of the structural unit cell (Figure S2).

together with the magnetic scattering from the CoO impurity,<sup>24</sup> was able to account completely for the additional intensity. The magnetic model has a propagation vector  $k = (1/2\ 0\ 0)$  and may be described in space group  $P_6ma2$  (28.94 in the BNS scheme; a setting of space group  $C_{6h}m'm2'$  (35.10.245) in the OG scheme) as described in Table 4 and Tables S8–S12. All

**Table 4. Summary of the Parameters Describing the Magnetic Symmetry and Unit Cell for BaCoSO**

magnetic space group	$P_6ma2$ (28.94) (BNS scheme)
$a$ (Å)	6.07894(6) ( $c$ in nuclear structure)
$b$ (Å)	12.6976(1)
$c$ (Å)	7.94770(8) ( $2 \times a$ in nuclear structure)
$V$ (Å <sup>3</sup> )	613.47(1)

four nearest neighbor couplings between Co<sup>2+</sup> ions via Co–O–Co and Co–S–Co linkages are antiferromagnetic, and the arrangement is unfrustrated. The refined moment on Co<sup>2+</sup> ions in BaCoSO was 2.75(2)  $\mu_B$ , which is similar to the range of 2.6–2.8  $\mu_B$  per Co<sup>2+</sup> ion estimated from the computation reported by Valldor et al.<sup>10</sup> and similar to the value in CaCoSO.

## CONCLUSIONS

We have described a new non-centrosymmetric insulating antiferromagnet CaCoSO in which the frustration between the moments of high spin Co<sup>2+</sup> ion spins inherent to the triangular arrangement is relieved by the adoption of a canted 120° structure. We have verified that the unfrustrated magnetic structure of the isostoichiometric compound BaCoSO, which has a different crystal structure, is that predicted from computation. In both compounds the ordered Co<sup>2+</sup> moments are similar at about 2.7  $\mu_B$ , reflecting a largely quenched orbital moment and a reduction below the spin-only value due to covalency.

## ASSOCIATED CONTENT

### Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.inorgchem.5b02615.

Tables of refined structural parameters (PDF)

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### Notes

The authors declare no competing financial interest.

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